INTERNATIONAL INDIAN SCHOOL BURAIDAH

CLASS- IX.

Worksheet

SUBJECT: CHEMISTRY

Ch#3 ATOMS AND MOLECULES

Q1. State law of conservation of mass.

Q2. State law of constant proportion taking ammonia as an example.

Q3. Which postulate of Dalton explains law of conservation of mass.

Q4.Give one example of diatomic liquid molecule.

Q5.State three main postulates of Dalton's Atomic theory.

Q6.What are the main limitations of Dalton's Atomic Theory.

Q7.Why was oxygen selected earlier for recording the relative atomic mass values?

Give reasons

Q8. Define a) molecular mass b) Formula unit mass c) Relative atomic mass d)

Atomicity

Q9. State the difference between Formula unit mass and molecular mass.

Q10. What is a polyatomic ion? Give two examples.

Q11. Name the cations and anions in the following :

NH_4Cl b) MgO c) Na_2SO_4 d) $CaCO_3$ e) $K_4[Fe(CN)_6]$ f) NaOH g) $AlCl_3$

h) $Cu(NO_3)_2$ j) $FeSO_3$ k) sodium acetate

Q12. Write the chemical formulae of the following compounds:

a) Iron (III) sulphate Aluminium sulphate

b) Iron(II) sulphate Silver Chloride

c) Copper (I) oxide Mercurous oxide

d) Copper(II) oxide Mercuric Chloride

e) Zinc Oxide Sodium phosphate

f) Ammonium nitrate Magnesium acetate

g) Aluminium oxide

Q13. What is a.m.u? How is it linked with relative atomic mass?

Q14. The Relative atomic mass of Boron is 9. What do you mean by this statement?

Q15. Name an element which is a) Tetra atomic b) Octa- atomic c) Triatomic

Q16.Calculate the formula unit mass of Aluminium sulphate.

(At wt of Al=27u e, S=32u, O=16u)

Q17. Which has more number of atoms? 10g of N_2 or 10g of NH_3

Q18.An element forms an oxide A_2O_3. What is the formula of the chloride of A? What is the valency of A?